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David Levy Inst. Ciencia de Materiales de Madrid, ICMM - CSIC

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The Sol-Gel Handbook: Synthesis, Characterization, and ...

A gel consists of a porous, three-dimensionally continuous solid network surrounding and supporting a continuous liquid phase (“ wet gel ”). In most sol – gel systems for the synthesis of oxide materials, gelation (i.e., formation of the gels) is due to the formation of covalent bonds between the sol parti-cles. Gel formation can be

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reversible when other bonds are involved, such as van der Waals forces or hydrogen bonds.

Part One Sol Gel Chemistry and Methods

Sol-gel synthesis of metal oxide can be done at relatively low temperature compared to the solid-state reactions. In general, sol-gel process involves formation of sol from homogeneously mixed...

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The sol – gel method involves two main reactions: (1) hydrolysis of the precursor in the acidic or basic mediums and (2) polycondensation of the hydrolyzed products. In this way a polymeric network is formed in which MNPs can be retained [126].
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This process can be summarized in six steps: (1) the formation of a stable metal precursor solution referred to as “ sol ” ; (2) the formation of a “ gel ” through a polycondensation reaction; (3) the

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Aging of the gel for hours or days, resulting in the expulsion of the solvent, i.e., Ostwald ripening, and the formation of a solid mass; (4) the drying of the gel of any liquids; (5) dehydration and surface stabilization; and (6) heat treatment of the gels at high temperatures to generate ...

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The Sol-Gel Handbook: Synthesis, Characterization, and ...

The diverse, international team of contributing authors of this reference clarify in extensive detail properties and applications of sol-gel science and technology as it pertains to the production of substances, active and non-active, including optical, electronic, chemical, sensor, bio- and structural materials.

Handbook of Sol-Gel Science and Technology - Processing ...

Alkoxides are ideal chemical precursors for sol – gel synthesis because they react readily with water. The reaction is called hydrolysis, because a hydroxyl ion becomes attached to the silicon atom as follows: $\text{Si (OR)}_4 + \text{H}_2\text{O} \rightarrow \text{HO – Si (OR)}_3 + \text{R – OH}$

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